

# Three-Dimensional Lithium-Ion Battery Model

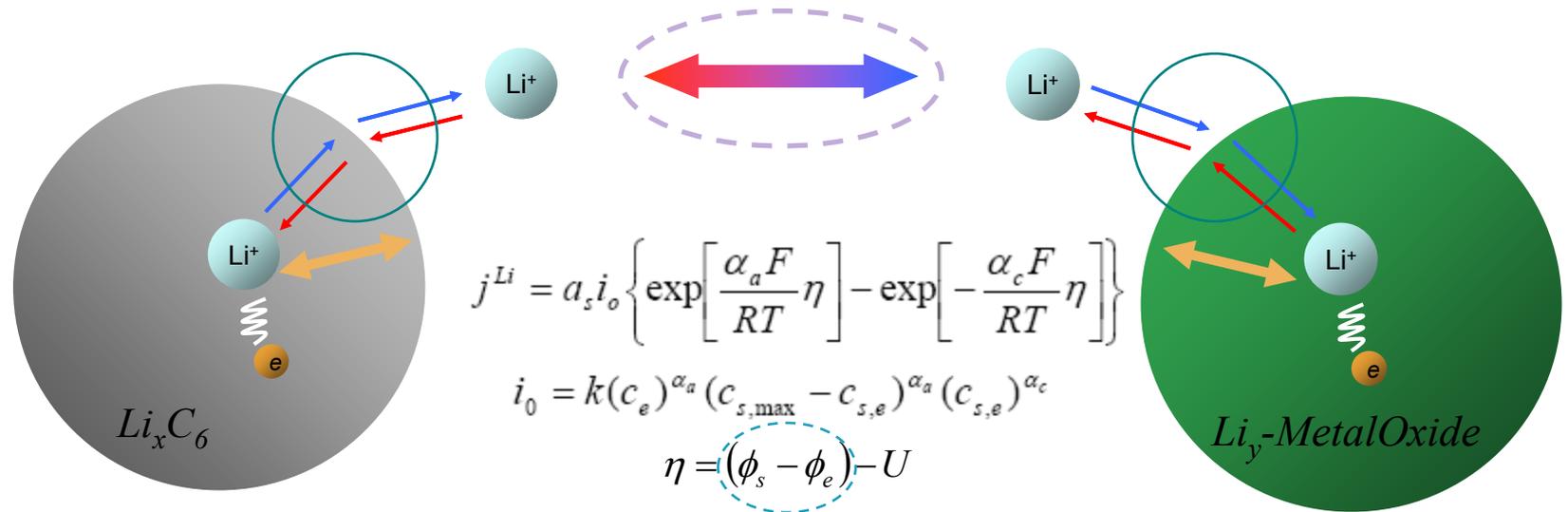
Understanding Spatial Variations in Battery Physics to Improve Cell Design, Operational Strategy, and Management



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# Multi-Scale Physics in Li-ion Battery



Electrochemical Kinetics

Solid-Phase Lithium Transport

Lithium Transport in Electrolyte

Charge Conservation/Transport

(Thermal) Energy Conservation

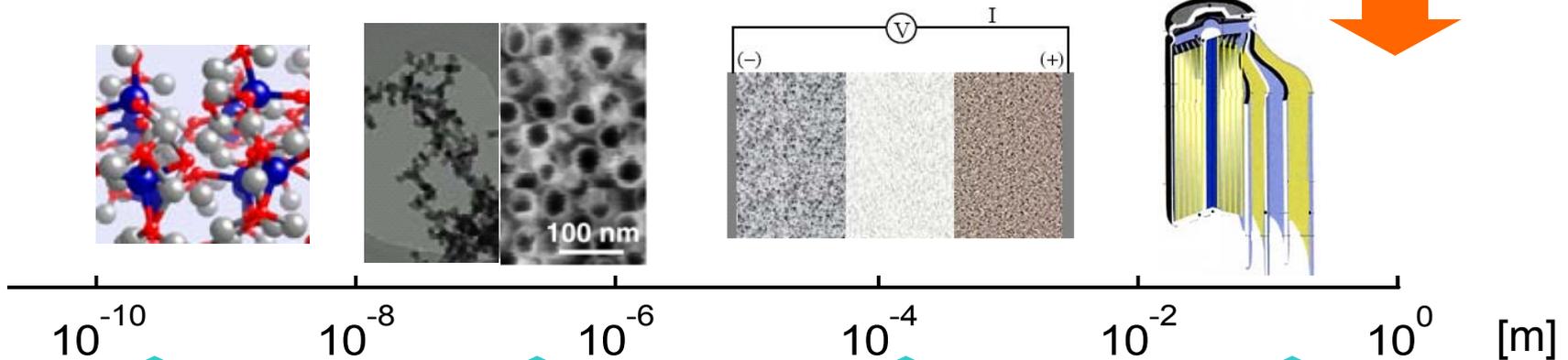
**Basic battery physics occurs in a wide range of length & time scales**

- Kinetics
- Phase transition
- Ion transport
- Energy dissipation
- Heat transfer

# Requirements & Resolutions

“**Requirements**” are usually defined in a macroscale domain and terms.

Performance  
Life  
Cost  
Safety



## Design of Materials

Voltage  
Capacity  
Lattice stability  
Kinetic barrier  
Transport property

## Design of Electrode Architecture

Li transport path (local)  
Electrode surface area  
Deformation & fatigue  
Structural stability  
Interface physics

## Design of Electrodes Pairing and Lithium Transport

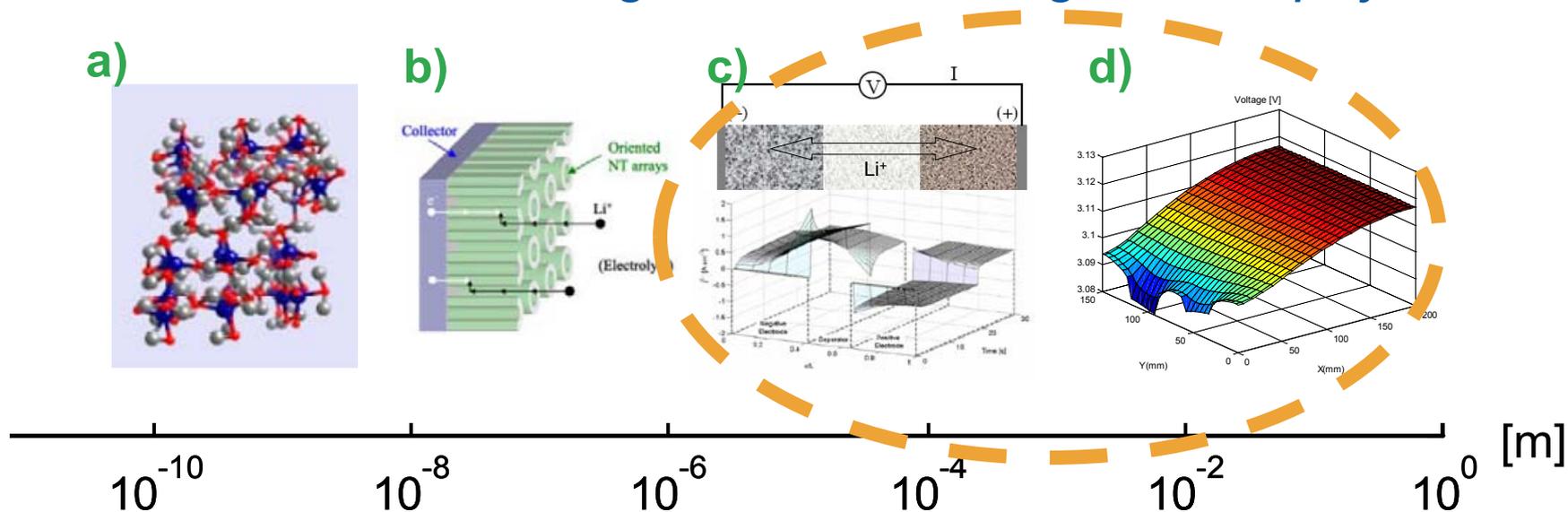
Electrodes selection  
Li transport  
Porosity, tortuosity  
Layer thicknesses  
Load conditions

## Design of Electronic Current & Heat Transport

Electric & thermal connections  
Dimensions, form factor  
Component shapes

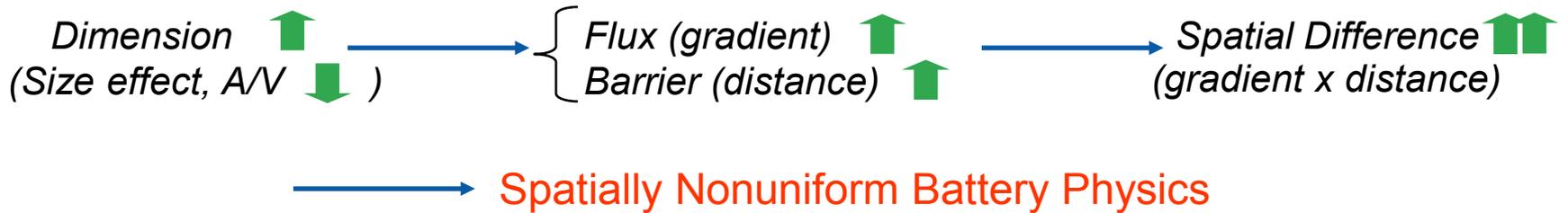
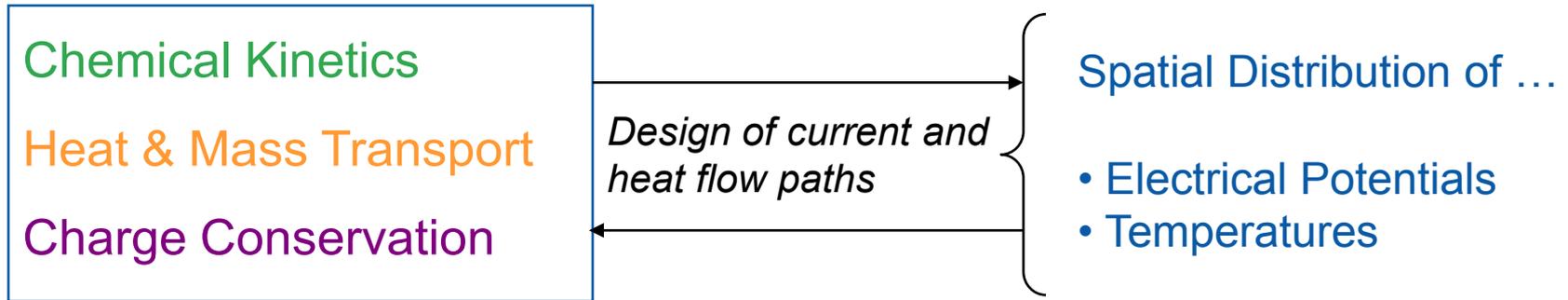
# NREL's Li-ion Battery Model Activities

*focusing on different length scale physics*



- a) Quantum mechanical and molecular dynamic modeling
- b) Numerical modeling resolving architecture of electrode materials
- c) Electrode-scale performance model
- d) Cell-dimension 3D performance model

# Why use a 3D model?



Enhanced understanding provides an opportunity for improving cell ...

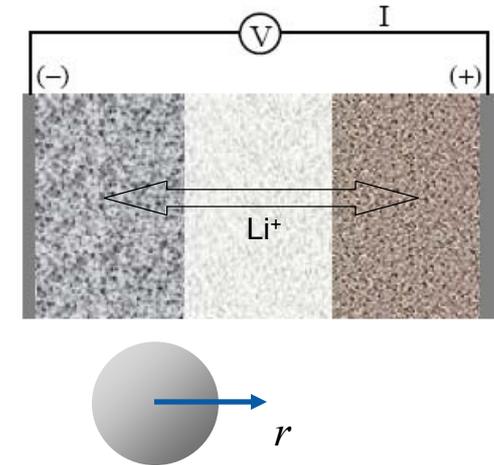
- *Design*
- *Operation Strategy*
- *Management*
- *Safety*

# Electrode-Scale Model

(Doyle, Fuller, and Newman, 1993)

✚ This model captures relevant *solid-state and electrolyte diffusion dynamics* and predicts the *current/voltage response* of a battery.

✚ Composite electrodes are modeled using *porous electrode theory*, meaning that the solid and electrolyte phases are treated as superimposed continua *without regard to microstructure*.



## Chemical Kinetics

$$j^{\text{Li}} = a_s i_o \left\{ \exp \left[ \frac{\alpha_a F}{RT} \eta \right] - \exp \left[ - \frac{\alpha_c F}{RT} \eta \right] \right\}$$

$$i_o = k (c_e)^{\alpha_a} (c_{s,\text{max}} - c_{s,e})^{\alpha_a} (c_{s,e})^{\alpha_c}$$

## NREL's Model

- Finite-Volume Method
- Matlab Environment

## Heat & Mass Transport

$$\frac{\partial c_s}{\partial t} = \frac{D_s}{r^2} \frac{\partial}{\partial r} \left( r^2 \frac{\partial c_s}{\partial r} \right)$$

$$\frac{\partial (\varepsilon_e c_e)}{\partial t} = \nabla \cdot (D_e^{\text{eff}} \nabla c_e) + \frac{1-t_+^o}{F} j^{\text{Li}} - \frac{\mathbf{i}_e \cdot \nabla t_+^o}{F}$$

$$\rho c_p \frac{\partial T}{\partial t} = \nabla \cdot (k \nabla T) + q'''$$

$$\nabla \cdot (\sigma^{\text{eff}} \nabla \phi_s) - j^{\text{Li}} = 0$$

$$\nabla \cdot (\kappa^{\text{eff}} \nabla \phi_e) + \nabla \cdot (\kappa_D^{\text{eff}} \nabla \ln c_e) + j^{\text{Li}} = 0$$

## Charge Conservation

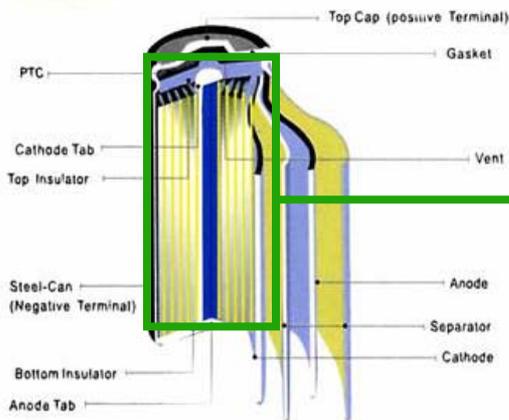


# Approach in the Present Study: Multi-Scale Multi-Dimensional (MSMD) Modeling

To Address ...

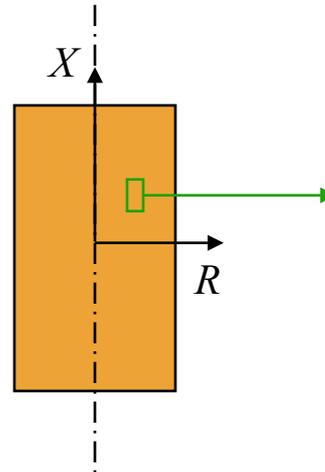
- Multi-scale physics from sub-micro-scale to battery-dimension-scales
- Difficulties in resolving microlayer structures in a computational grid

**Simulation Domain**



=

**Macro Grid**



+

**Micro Grid**

(Grid for Sub-grid-scale Model)

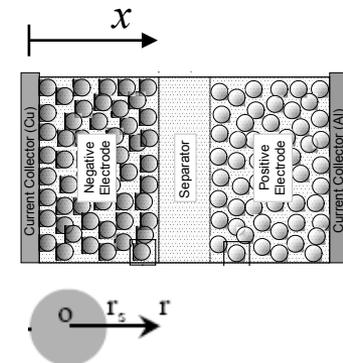


Image source: [www.dimec.unisa.it](http://www.dimec.unisa.it)

# Solution Variables

## Detailed Structure

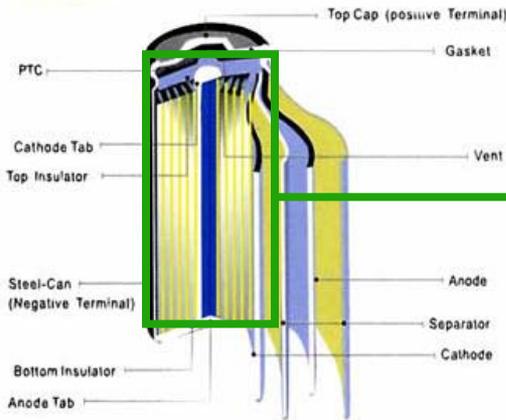
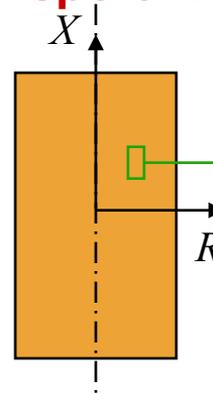
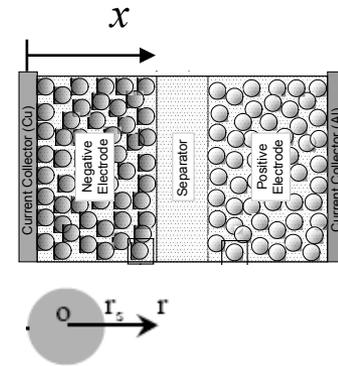


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## Cell Dimension Transport Model

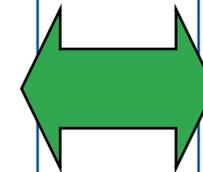


## Electrode Scale Submodel (1D)



**NOTE:**  
Selection of the “sub-grid electrochemical model” is independent of the “macro-grid model” selection.

$$\begin{aligned}
 &T(X, R, t) \\
 &V(X, R, t) \\
 &i(X, R, t) \\
 &SOC(X, R, t) \\
 &Q(X, R, t) = \int_x Q_i \frac{A dx}{V}
 \end{aligned}$$



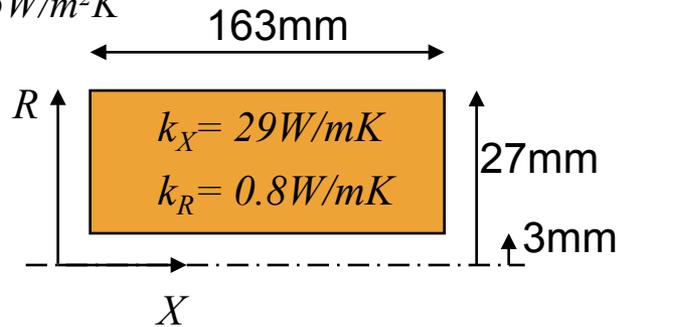
$$\begin{aligned}
 &\phi_s(X, R, x, t) \\
 &\phi_e(X, R, x, t) \\
 &c_s(X, R, x, r, t) \\
 &c_e(X, R, x, t) \\
 &j_{Li}(X, R, x, t) \\
 &Q_i(X, R, x, t)
 \end{aligned}$$

# Model Combination

**Axisymmetric FVM Model for Macro-Domain Model  
+ 1D FVM Model for Electrochemistry Submodel**

## 150A Constant Current Discharge

$T_{\infty} = 35^{\circ}\text{C}$   
 $h = 25\text{W/m}^2\text{K}$

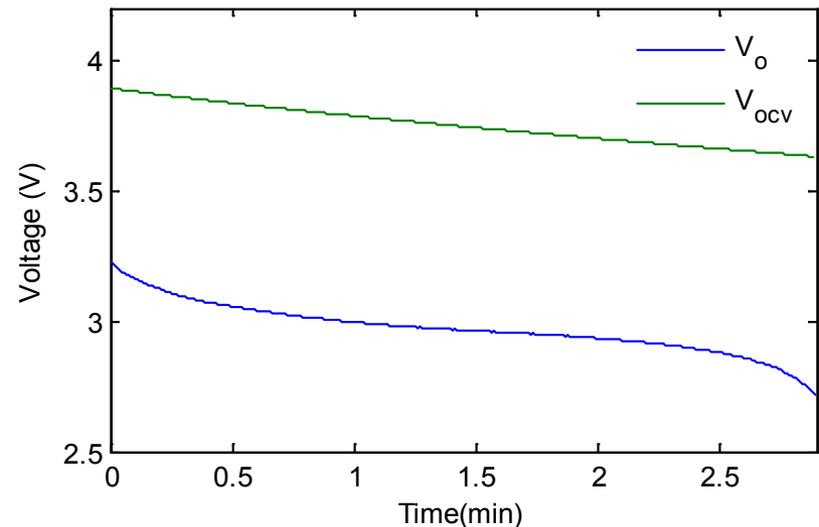


Grid#	X	R	x	r
	12	8	14	4

time step size: 0.5 sec, time step #: 360

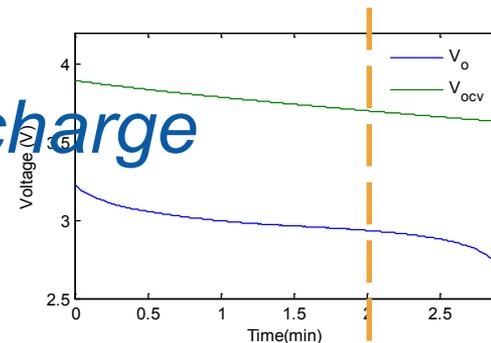
Simulation time: 3 minutes

**Computation time: 98 minutes (Windows/PC)**

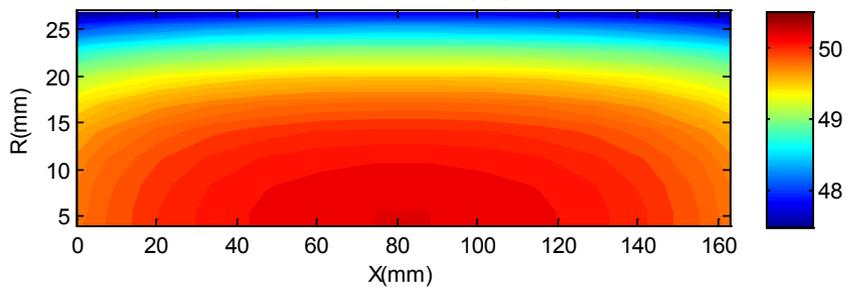


# Simulation Results

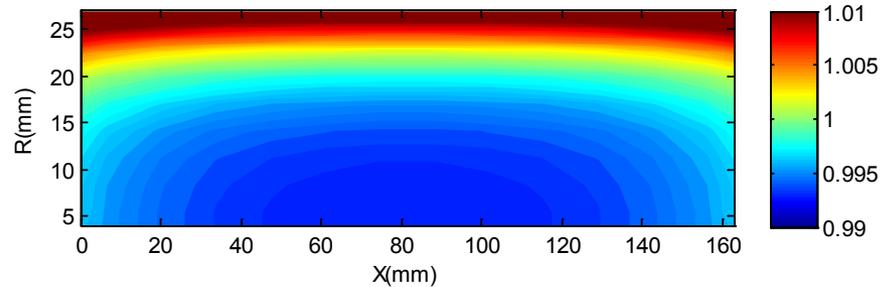
*Snapshot 2 minutes after start of discharge*



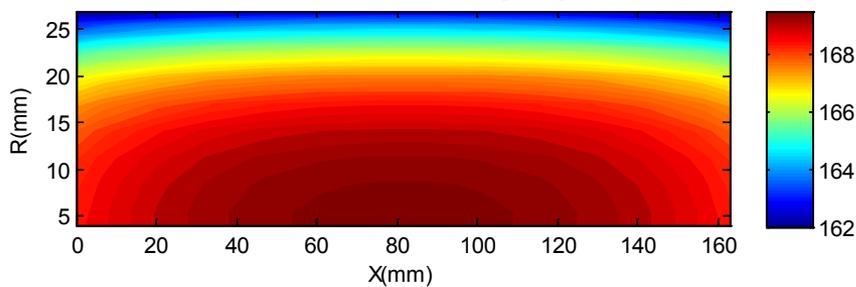
TEMPERATURE [ $^{\circ}\text{C}$ ]



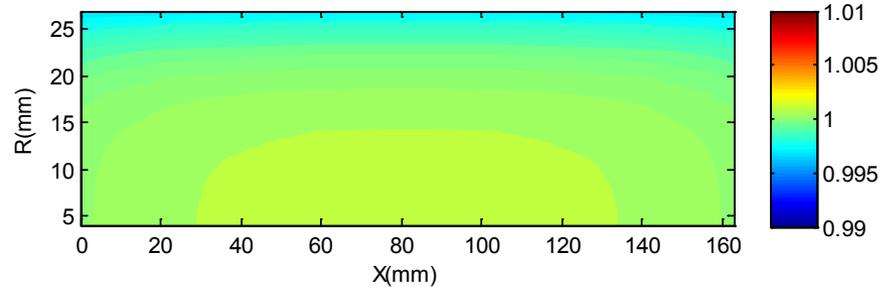
Normalized ANODE SURFACE CONCENTRATION



CURRENT PRODUCTION [ $\text{A}/\text{m}^2$ ]



Normalized CHATHODE SURFACE CONCENTRATION



# Another Combination Choice

*Axisymmetric FVM Model for Macro-Domain Model  
+ State Variable Model (SVM) for Submodel*

MSMD model incorporating SVM Submodel runs ~1.75 faster than real time.

## ***SVM is preferred because of its fast execution***

- ❑ SVM, developed by Kandler Smith (NREL), quickly solves “Newman type” governing equations using numerical schemes for calculating load reduction.
- ❑ Dropping very fast battery responses (approx. 60 Hz or more) is one of the main calculation order reduction methods used in the model.
- ❑ SVM is **promising for use in on-board BMS reference model** because of its fast execution and capability to provide nonmeasurable electrochemical parameters and current and voltage responses with potentially better accuracy.

### ***For details about the State Variable Model:***

See the Poster Presentation by Kandler Smith (NREL) titled,  
“Fast Running Electrochemistry-Based Models for Battery Design, Integration, and Control”

# Analysis

## Temperature Variation in a Cylindrical Cell

- *Uniform Potential Assumption*
- *Impact of Aspect Ratio*
- *Impact of Cell Size*

## Temperature & Potential Variation in a Prismatic Cell

- *Impact of Tab Location and Size*

# Analysis

## Temperature Variation in a Cylindrical Cell

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## Temperature & Potential Variation in a Prismatic Cell

- *Impact of Tab Location and Size*

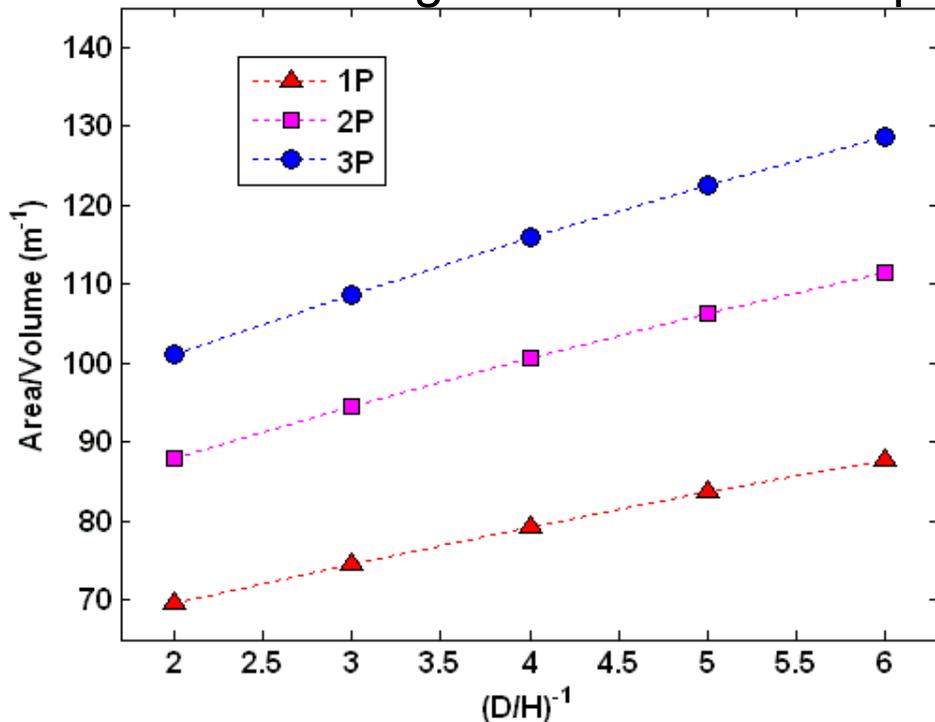
# Considerations for Addressing Thermal Issues in PHEV-type Cells

- ❏ High energy *and* high power requirements
- ❏ Large format may be preferred to small cells
  - Fewer number of components
  - Fewer interconnects
  - Less monitoring & balancing circuitry
  - Less expensive
  - Less weight
- ❏ Significant heating may be possible, depending on power profile
- ❏ Internal temperature imbalance can lead to unexpected performance and life degradation

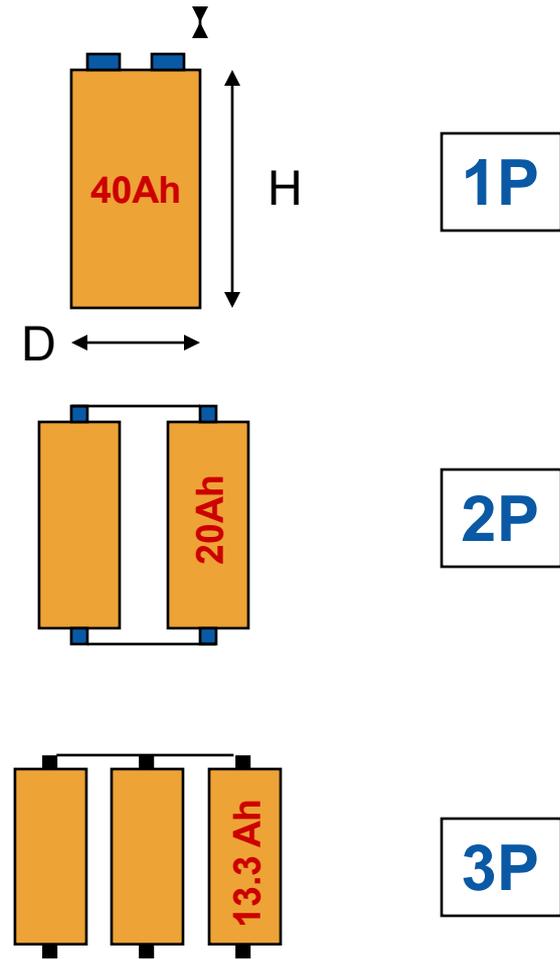
# Analysis Parameters

For a fixed capacity (electrode volume), surface area for heat rejection can be increased via:

- Reducing D/H ratio
- Increasing number of cells in parallel (#P)



\*Surface area includes side, top & bottom of can. All cells assumed to have inactive inner mandrel with 8mm diameter.

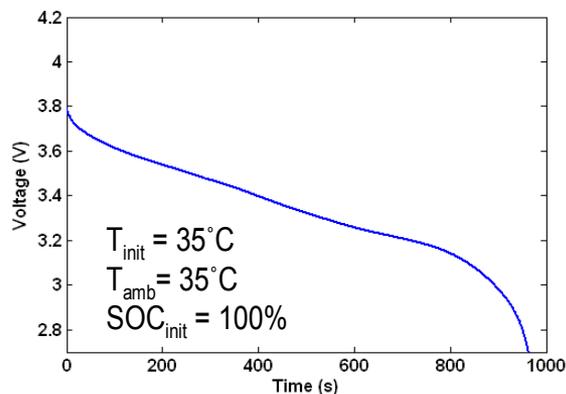


# Two Usage Profiles

*The two cases explored in this presentation:*

## 1 150A Max. Cont. Discharge

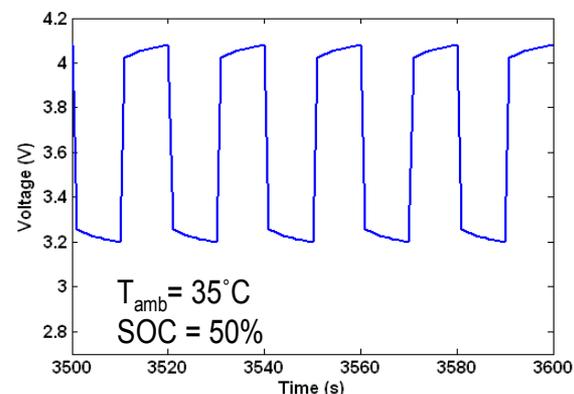
- Transient, Charge Depleting
- Air Convection ( $15 \text{ W/m}^2\text{K}$ )



→ Moderate Thermal Condition

## 2 200A Geometric Cycle

- Steady-State, Charge Sustaining
- Liquid Cooling ( $150 \text{ W/m}^2\text{K}$ )



→ Severe Thermal Condition

# Results: 150 A Continuous Discharge

## Transient Results

$$D/H = 1/4$$

$$h = 15 \text{ W/m}^2\text{K}$$

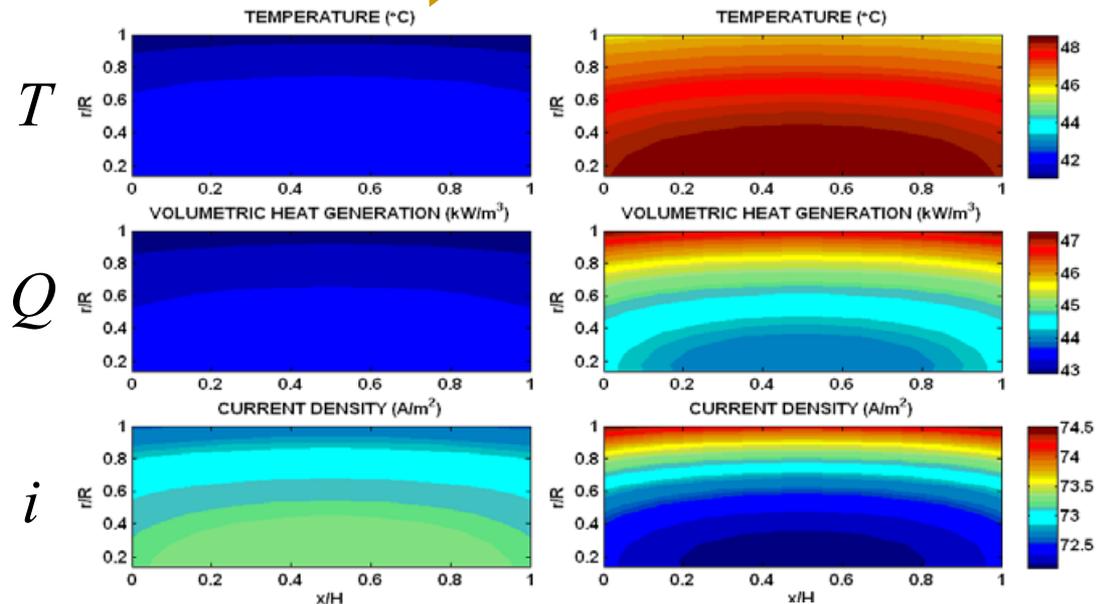
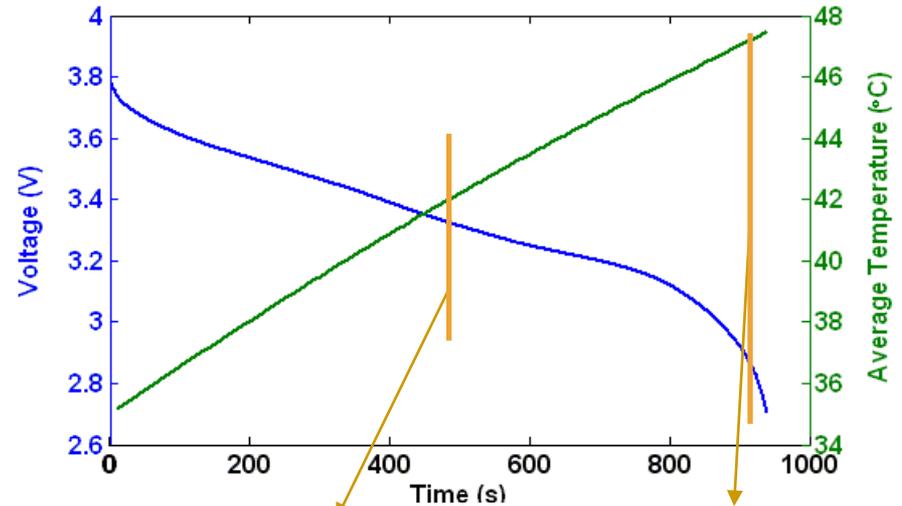
$$T_{\text{amb}} = 35^\circ\text{C}$$

After 500 seconds of discharge:

- Cell center is slightly warmer than exterior
- Preferential reaction current at cell center

Near the end of discharge:

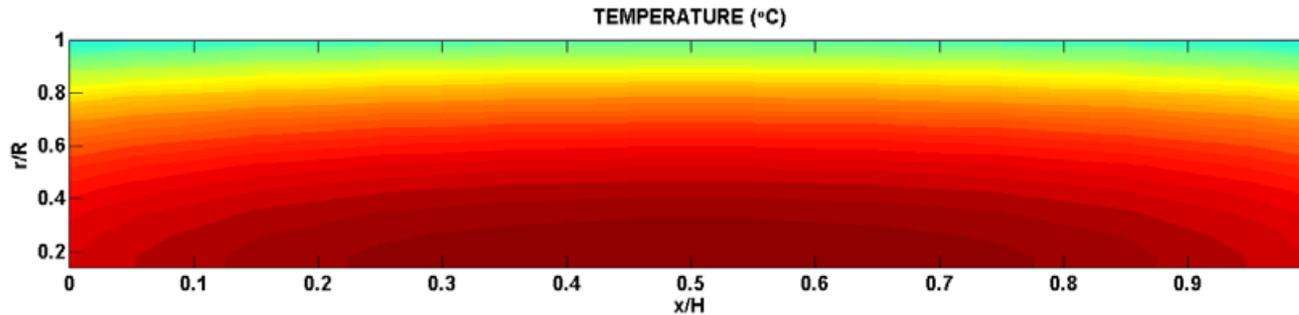
- Cell center depleted/saturated
- Preferential reaction current at cell exterior



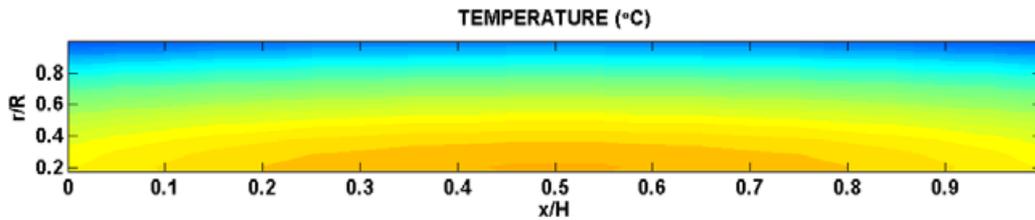
# 150 A Single Discharge (at End)

$D/H = 1/4$   
 $h = 15 \text{ W/m}^2\text{K}$   
 $T_{\text{amb}} = 35 \text{ C}$

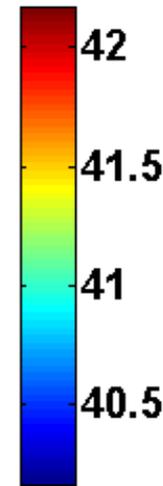
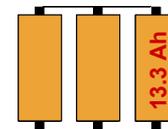
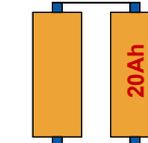
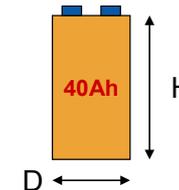
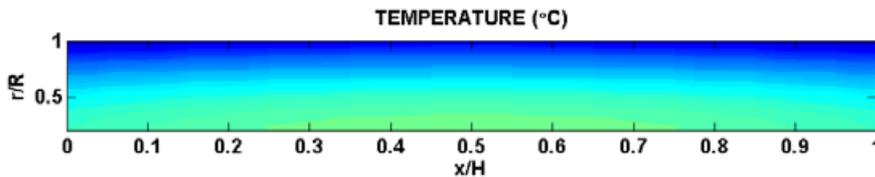
1P



2P



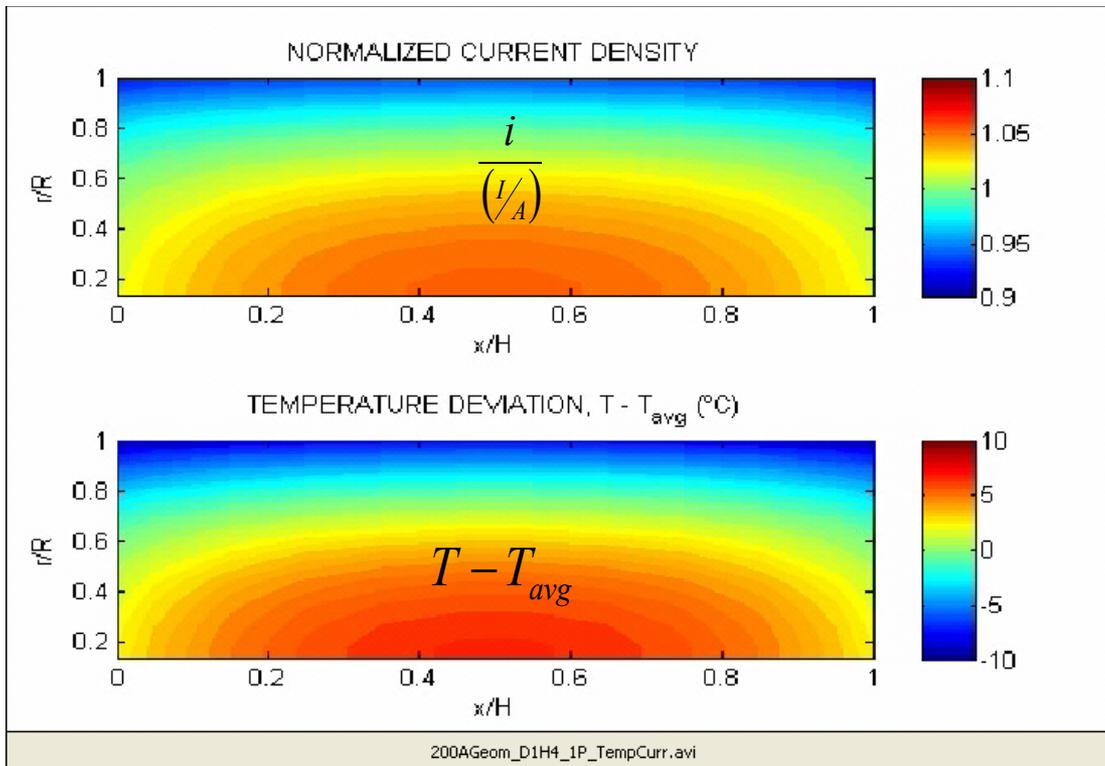
3P



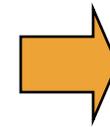
*Moderate usage + air convection = small internal gradients*

# 200 A Geometric Cycling

## At Steady State



~16% difference in local current production



$\Delta T = \sim 17$  °C

$D/H = 1/4$

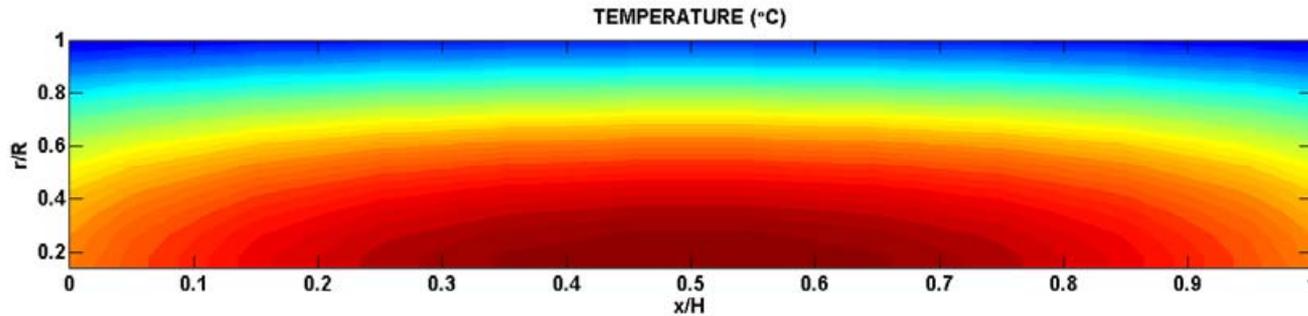
$h = 150$  W/m<sup>2</sup>K

$T_{amb} = 35$  °C

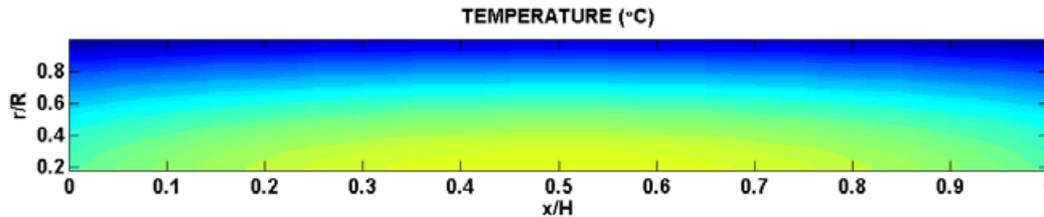
1P size cell (40 Ah)

# 200 A Geometric Cycle (Steady-State)

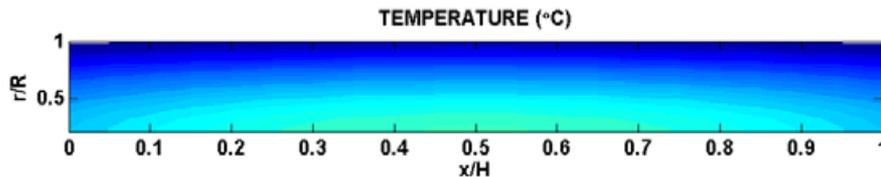
1P



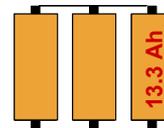
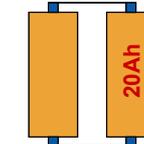
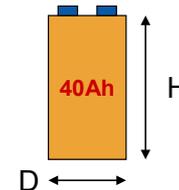
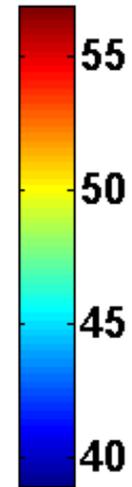
2P



3P



$D/H = 1/4$   
 $h = 150 \text{ W/m}^2\text{K}$   
 $T_{\text{amb}} = 35 \text{ C}$



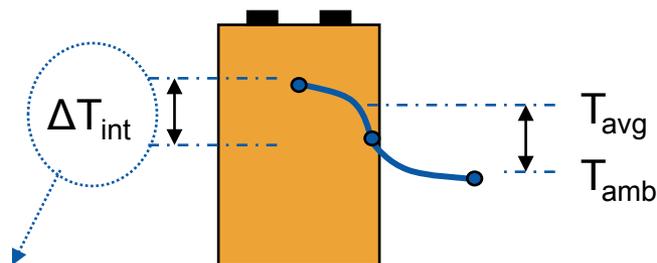
*Severe usage + liquid cooling = large internal gradients*

# 200 A Geometric Cycle (Steady-State)

## Internal Temperature Difference

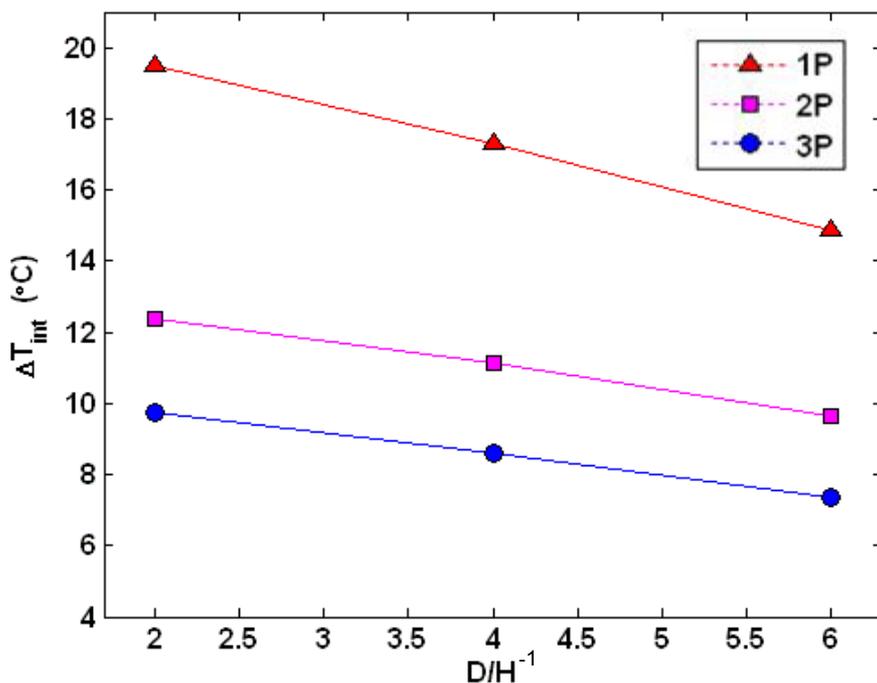
D/H ratio  $\sim 4.5^\circ\text{C}$   $\updownarrow$

2P  $\sim 6.0^\circ\text{C}$ , 3P  $\sim 9.0^\circ\text{C}$   $\downarrow$



$h = 150 \text{ W/m}^2\text{K}$

$T_{\text{amb}} = 35^\circ\text{C}$



- Under severe usage, low  $D/H$  and/or  $>1P$  designs significantly reduce thermal stress
- Larger diameter leads to higher internal gradient
- Multidimensional electrochemical cell model quantified the *impacts of  $D/H$  aspect ratio and cell size on the internal temperature difference.*

# Analysis

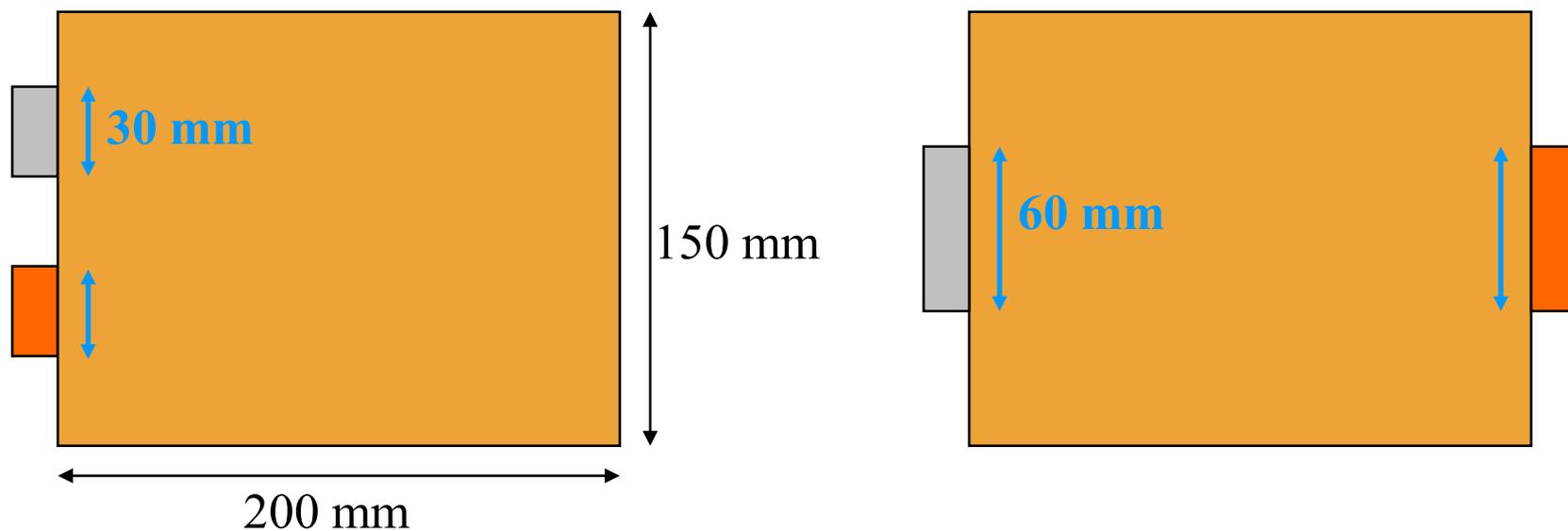
## Temperature Variation in a Cylindrical Cell

- *Uniform Potential Assumption*
- *Impact of Aspect Ratio*
- *Impact of Cell Size*

## **Temperature & Potential Variation in a Prismatic Cell**

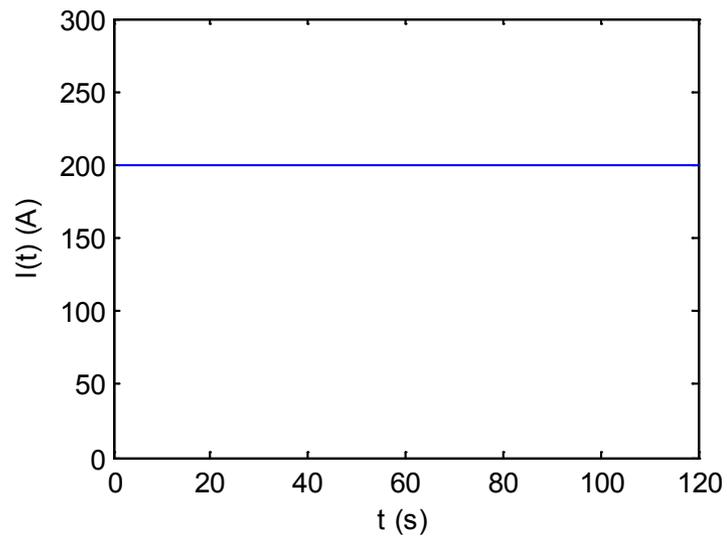
- *Impact of Tab Location and Size*

## Impact of Tab Location & Size

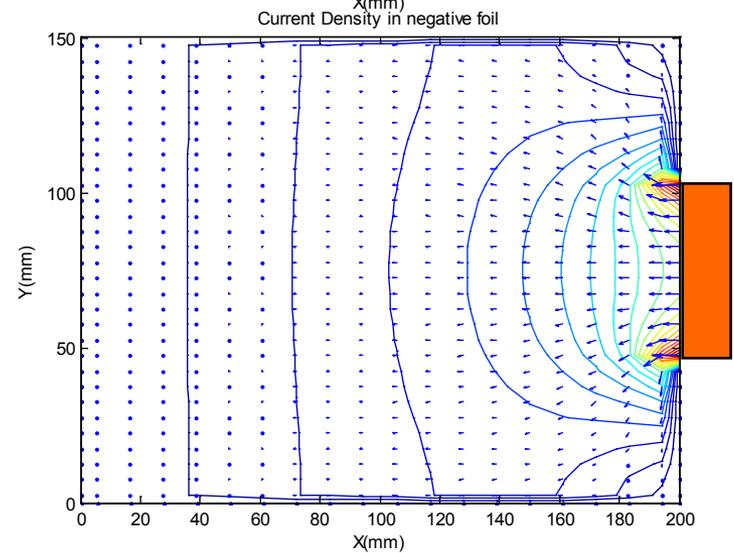
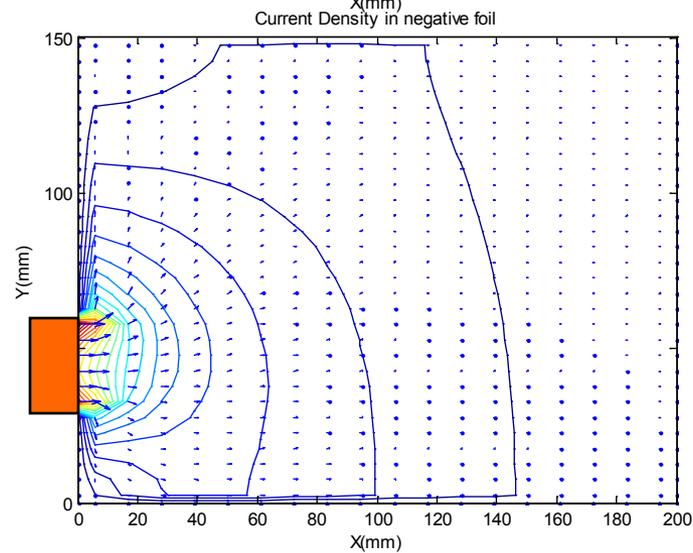
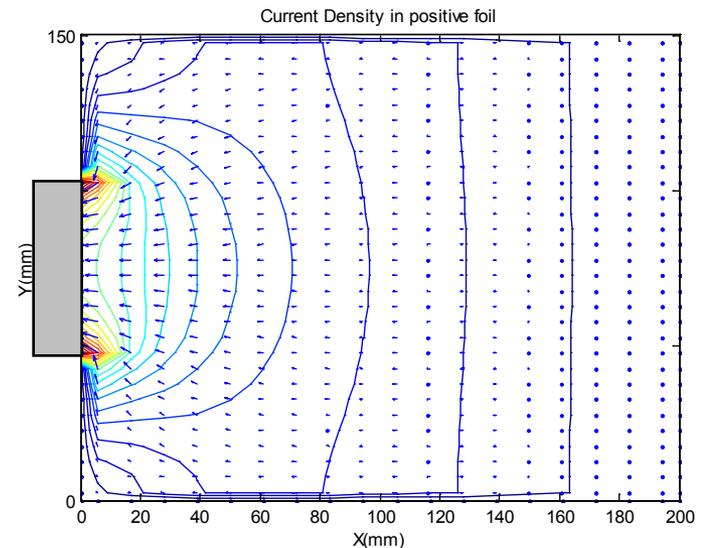
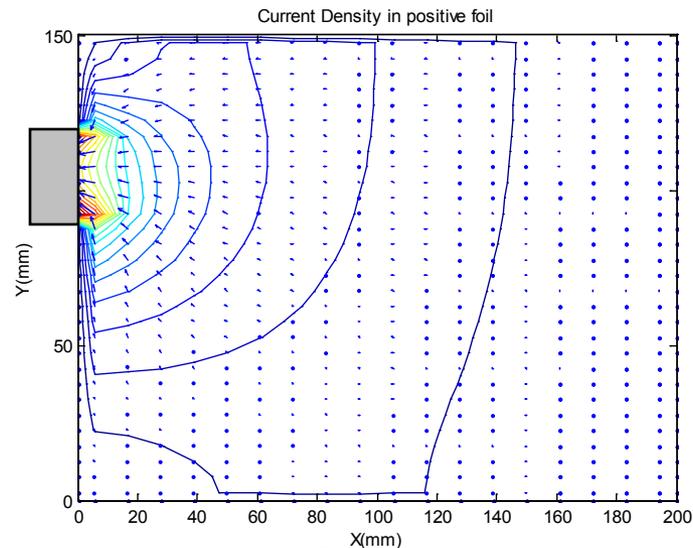


- Thickness: 12 mm
- 40 Ah
- 2-minute discharge, 200 A
- 200A geometric cycle

# 200A Discharge for 2 minutes

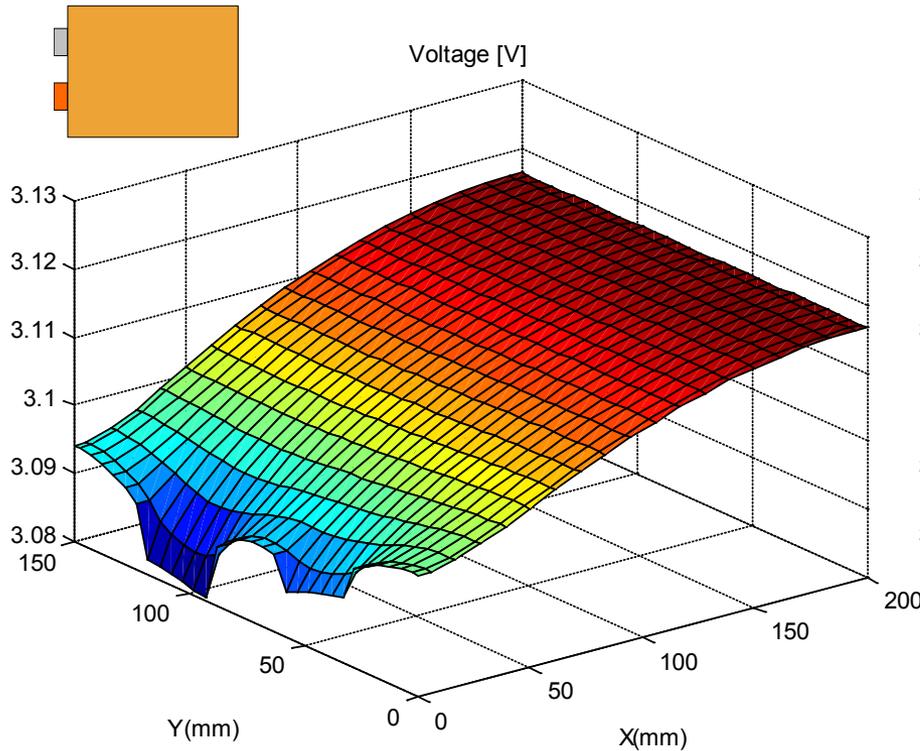


# Current Field – 2-min 200 A discharge

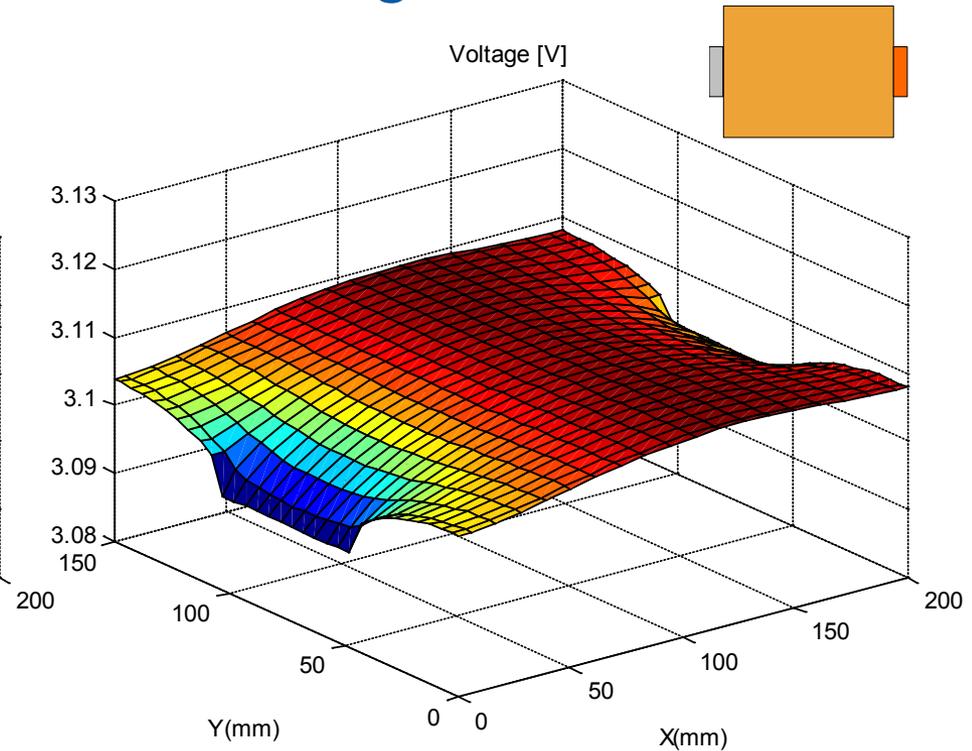


# Voltage across Current Collector Foils

– 2-min 200 A discharge

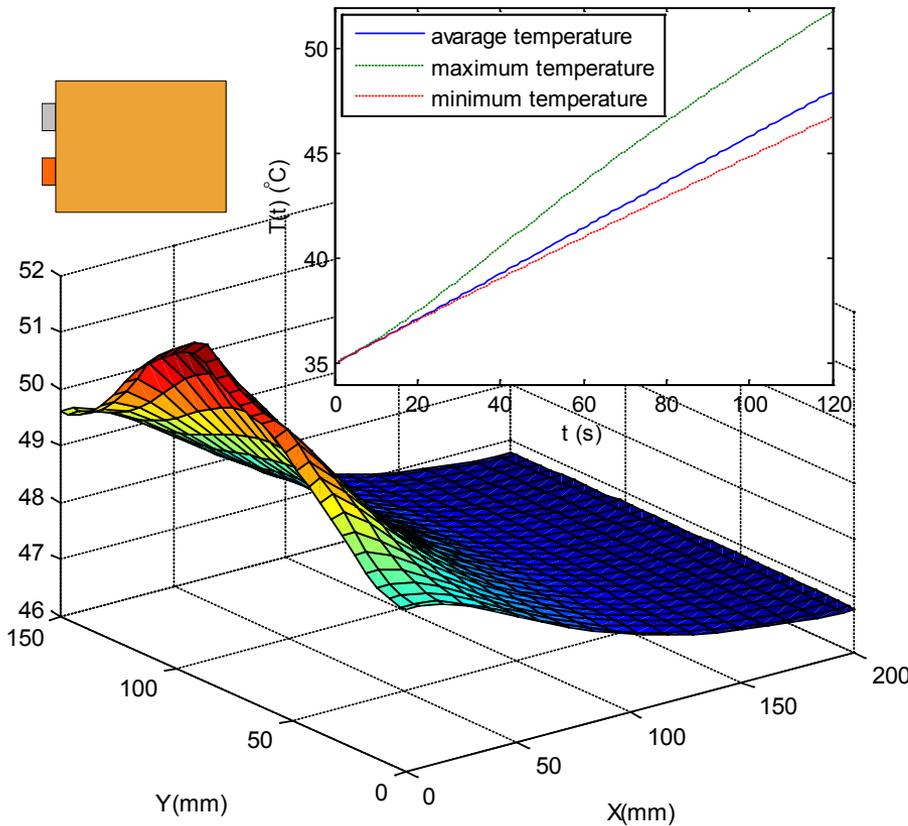


$$V_{\max} - V_{\min} = 0.0364 \text{ V}$$

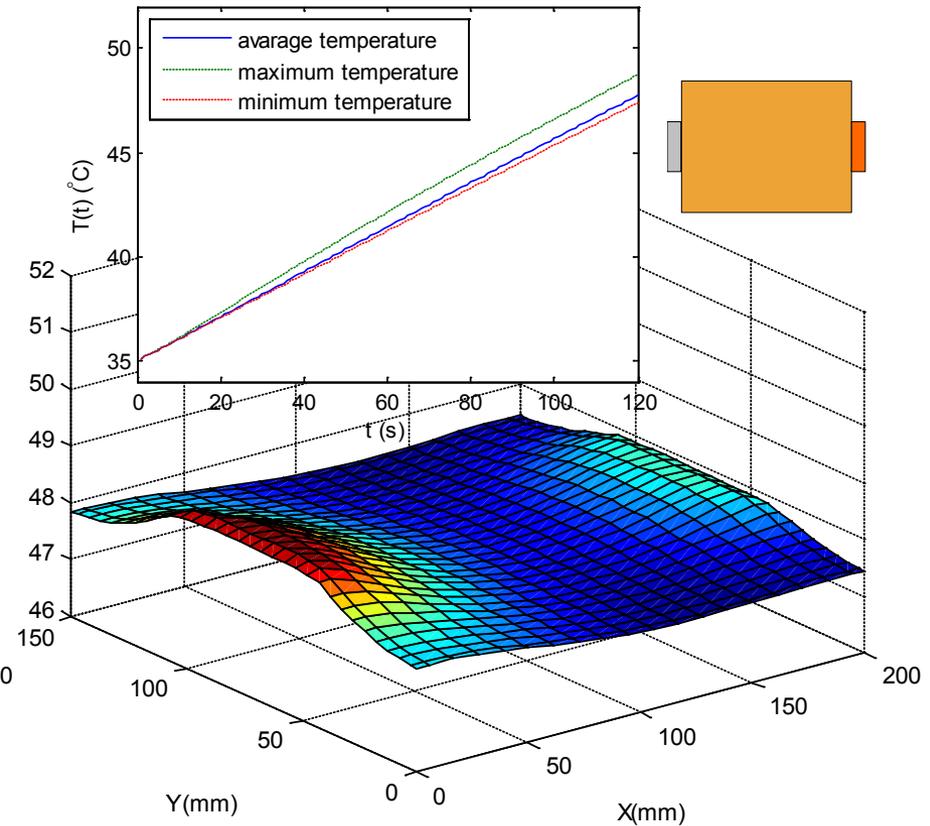


$$V_{\max} - V_{\min} = 0.0154 \text{ V}$$

# Temperature – 2-min 200 A discharge

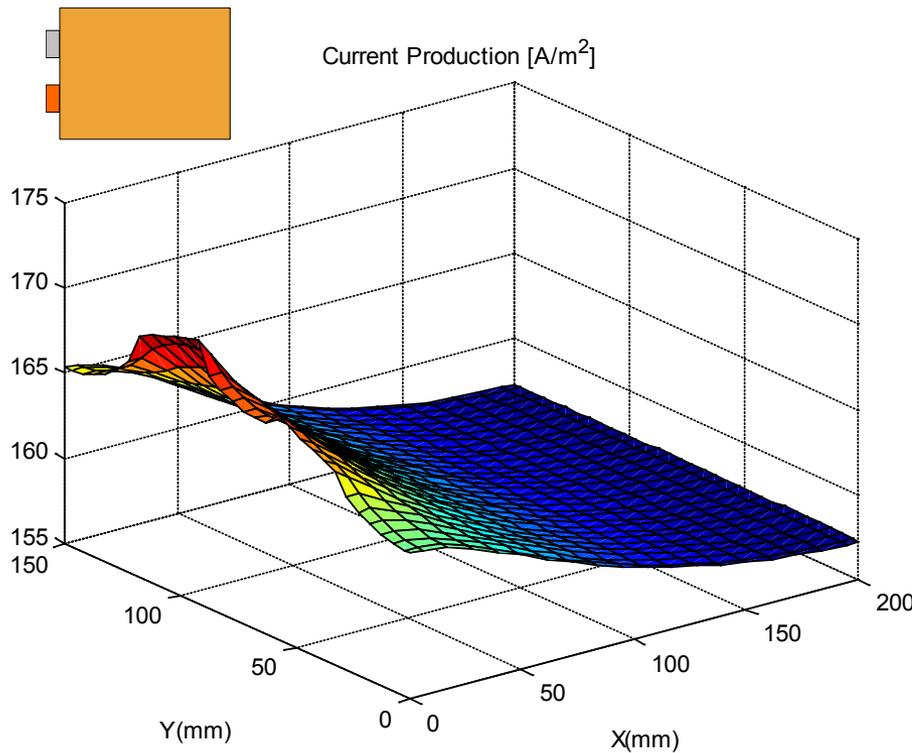


$$T_{\max} - T_{\min} = 5.03^{\circ}\text{C}$$

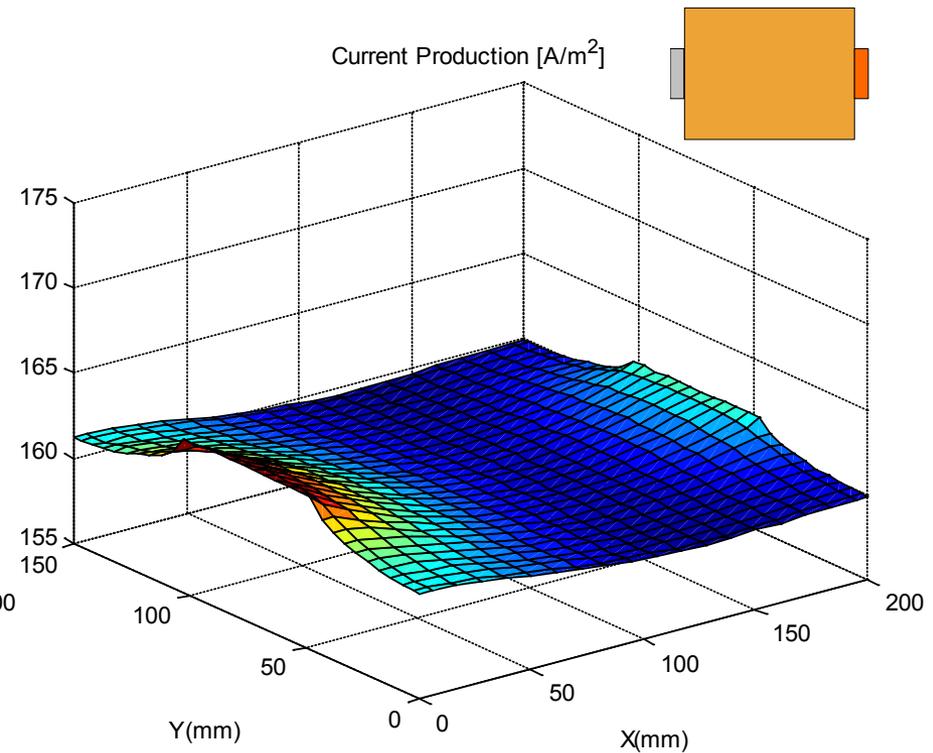


$$T_{\max} - T_{\min} = 1.35^{\circ}\text{C}$$

# Current Production – 2-min 200 A discharge

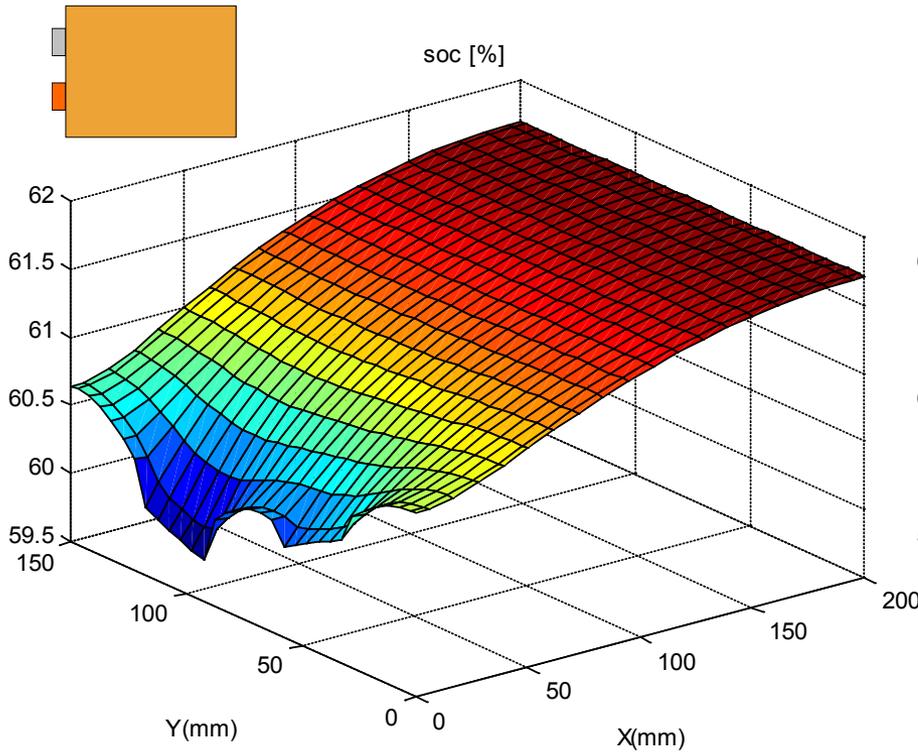


$$i_{\max} - i_{\min} = 13.2 \text{ A/m}^2$$

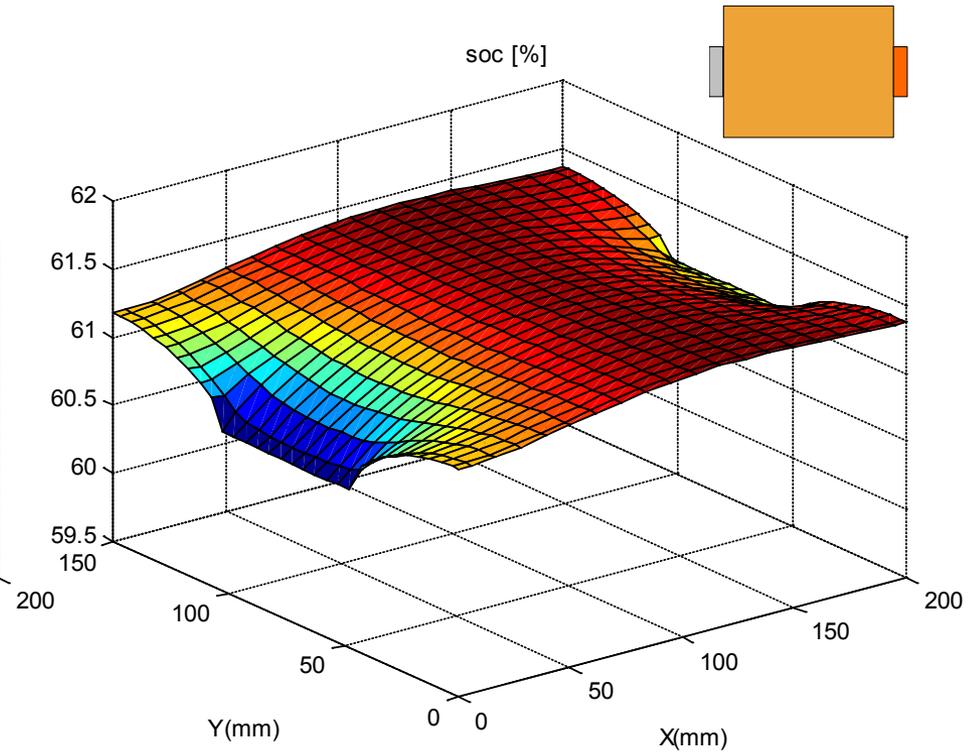


$$i_{\max} - i_{\min} = 4.54 \text{ A/m}^2$$

# SOC – 2-min 200 A discharge

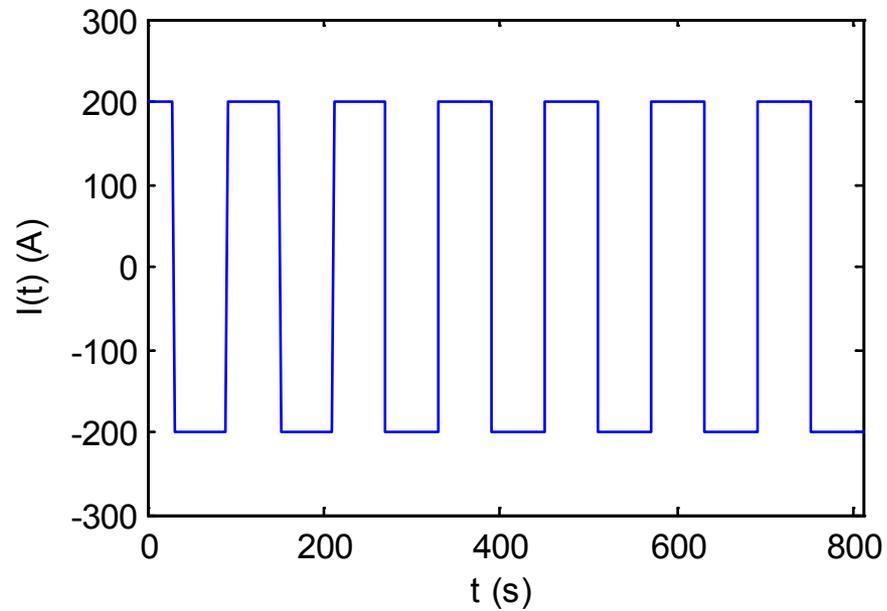


$$\text{SOC}_{\max} - \text{SOC}_{\min} = 1.91\%$$

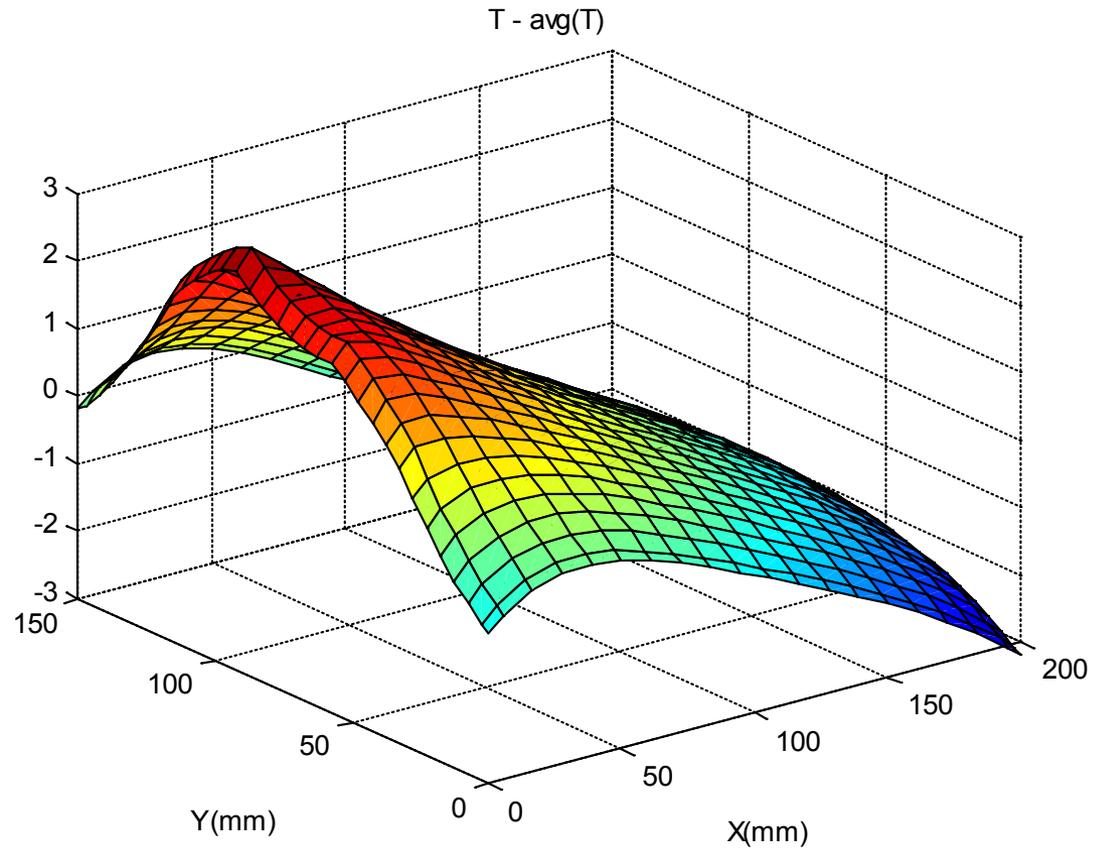
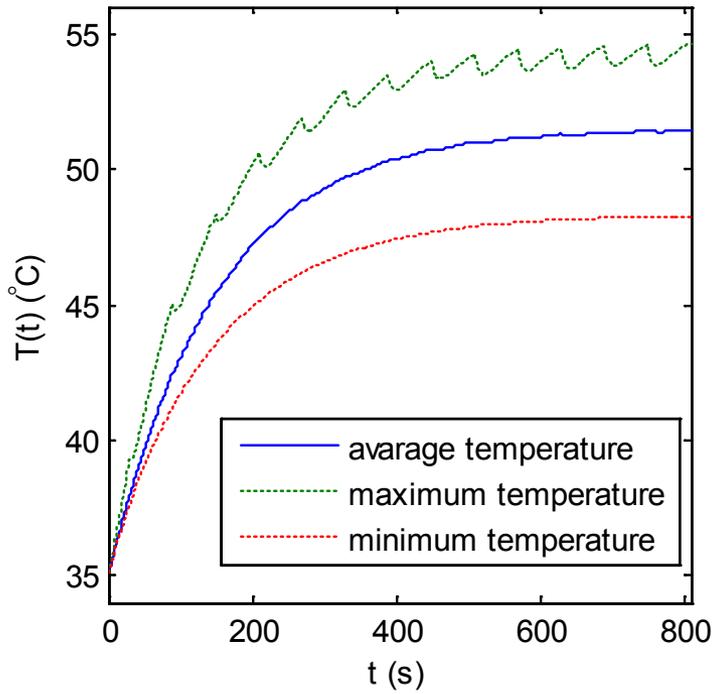


$$\text{SOC}_{\max} - \text{SOC}_{\min} = 0.76\%$$

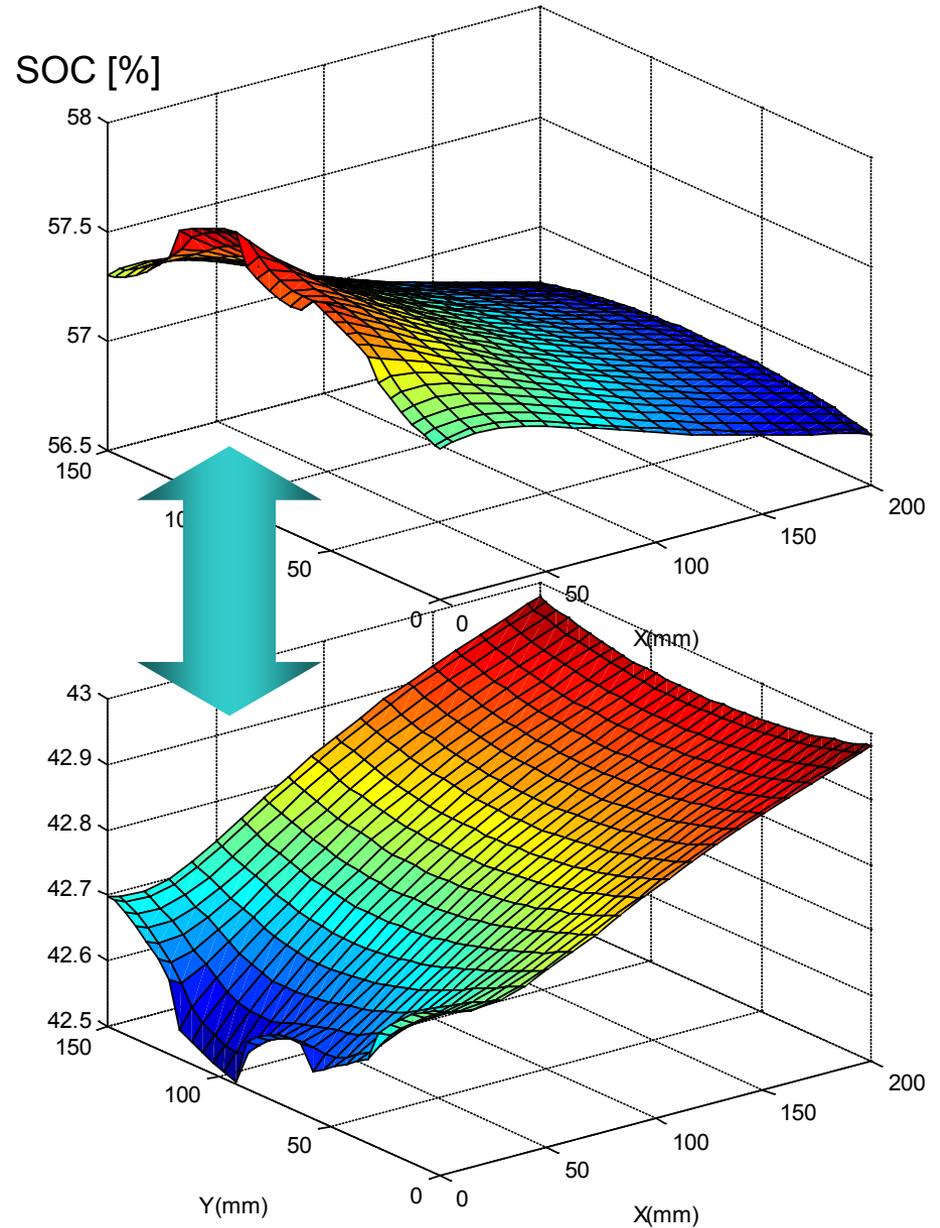
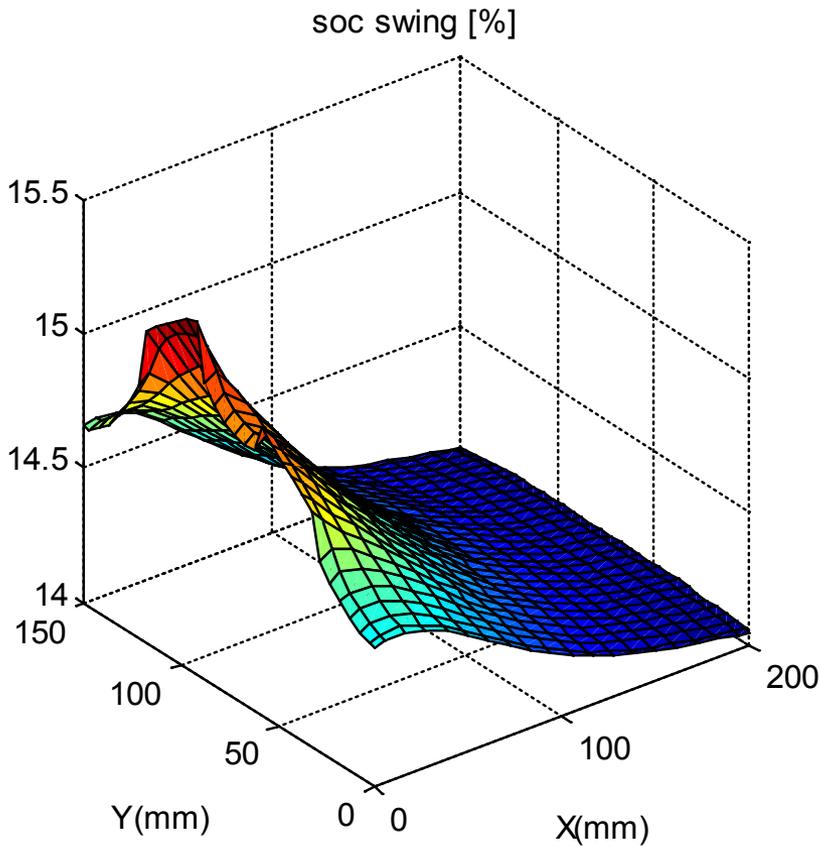
# 200A Geometric Cycling



# Temperature Variation



# SOC swing



# Summary

❑ **Nonuniform battery physics**, which is more probable in large-format cells, can cause unexpected performance and life degradations in lithium-ion batteries.

❑ **A three-dimensional cell performance model** was developed by integrating an electrode-scale submodel using a **multiscale modeling** scheme.

❑ The developed tool will be used to provide better understanding and help answer **engineering questions** about improving *cell design*, *cell operational strategy*, *cell management*, and *cell safety*.

❑ Engineering Questions to be addressed in *future works* include ...

*What is the optimum form-factor and size of a cell?*

*Where are good locations for tabs or current collectors?*

*How different are measured parameters from their non-measurable internal values?*

*Where is the effective place for cooling? What should the heat-rejection rate be?*

*How does the design of thermal and electrical paths impact under current-related safety events, such as internal/external short and overcharge?*

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